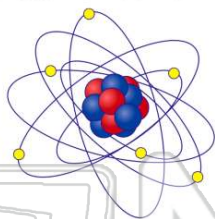




Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

- 1 Which **type of bond** exists between an atom of carbon and an atom of fluorine?

A ionic  
 B metallic  
 C polar covalent  
 D nonpolar covalent



- 2 The **oxidation number** of a **reducing agent** can change from

A -1 to -3  
 B -2 to -1  
 C 3 to -1  
 D 4 to -3

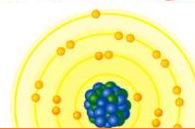
- 3 As an atom of **nitrogen** gains electrons, its **oxidation number**

A decreases  
 B increases  
 C remains the same



- 4 What is the total number of occupied sublevels in the **third principal energy level** of a **zinc atom** in the ground state?

A 1  
 B 2  
 C 3



## PREVIEW

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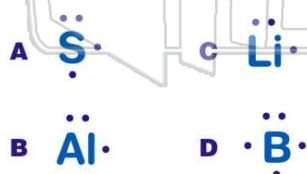
- 7

A 1  
 B 2  
 C 3  
 D 4



A a smaller number of electrons  
 B a smaller number of electrons  
 C an electron with greater energy  
 D an electron with less energy

- 9 Which **electron-dot symbol** correctly represents an atom of its given element?



- 10 Which statement best describes the substance that results when electrons are **transferred from a metal to a nonmetal**?

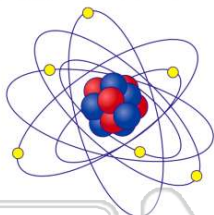
A It contains ionic bonds and has a low melting point.  
 B It contains ionic bonds and has a high melting point.  
 C It contains covalent bonds and has a low melting point.  
 D It contains covalent bonds and has a high melting point.



## ANSWER KEY

Which **type of bond** exists between an atom of carbon and an atom of fluorine?

- A ionic
- B metallic
- C polar covalent
- D nonpolar covalent



(c)

The **oxidation number** of a **reducing agent** can change from

- A -1 to -3
- B -2 to -1
- C 3 to -1
- D 4 to -3

(b)

As an atom of **nitrogen** gains electrons, its **oxidation number**

- A decreases
- B increases
- C remains the same

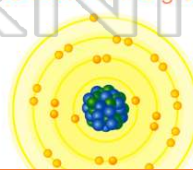


(a)

What is the total number of occupied sublevels in the **third principal energy level** of a **zinc atom** in the ground state?

- A 1
- B 2
- C 3
- D 4

(c)



## PREVIEW

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- C 3
- D 4



- C an electron with greater energy
- D an electron with less energy

Which **electron-dot symbol** correctly represents an atom of its given element?



(b)

Which statement best describes the substance that results when electrons are **transferred from a metal to a nonmetal**?

- A It contains ionic bonds and has a low melting point.
- B It contains ionic bonds and has a high melting point.
- C It contains covalent bonds and has a low melting point.
- D It contains covalent bonds and has a high melting point.

(b)