



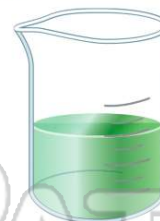
Name _____ Class _____ Date _____

1 The **name** of the compound **KClO₂** is **potassium**

- A hypochlorite
- B chlorite
- C chlorate
- D perchlorate

2 Which **compound** forms a **green aqueous solution**?

- A RbCl
- B CaCl₂
- C NiCl₂
- D ZnCl₂

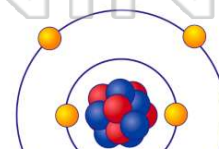


3 The **percent by mass of carbon** in **HC₂H₃O₂** is equal to

- A $\frac{12}{60} \times 100$
- B $\frac{24}{60} \times 100$
- C $\frac{60}{24} \times 100$
- D $\frac{60}{12} \times 100$

4 In which compound does **carbon** have an **oxidation state of -4**?

- A CO
- B CO₂
- C CCl₄
- D CH₄



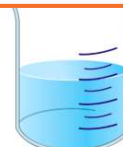
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- A Cl_2 and O_3
- B HO and H_2O
- C NO_2 and N_2O_4

- A NH_4Cl
- B NaNO_3
- C KCl



9 Which **compound** is a **nonelectrolyte**?

- A HNO₃
- B H₂SO₄
- C NaOH
- D CH₃OH



10 The **chemical formula** for **nickel (II) bromide** is

- A Ni₂Br
- B NiBr₂
- C N₂Br
- D NBr₂



ANSWER KEY

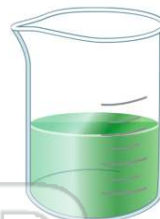
The **name** of the compound KClO_2 is **potassium**

- A hypochlorite
- B chlorite
- C chlorate
- D perchlorate

(b)

Which **compound** forms a **green aqueous solution**?

- A RbCl
- B CaCl_2
- C NiCl_2
- D ZnCl_2



(c)

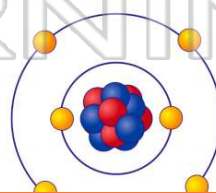
The **percent by mass of carbon** in $\text{HC}_2\text{H}_3\text{O}_2$ is equal to

- A $\frac{12}{60} \times 100$
- B $\frac{24}{60} \times 100$
- C $\frac{60}{24} \times 100$
- D $\frac{60}{12} \times 100$

(b)

In which compound does **carbon** have an **oxidation state of -4**?

- A CO
- B CO_2
- C CCl_4
- D CH_4



(d)

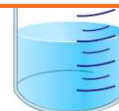


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D NO_2 and N_2O_4

- C NaNO_3
- D KCl



Which **compound** is a **nonelectrolyte**?

- A HNO_3
- B H_2SO_4
- C NaOH
- D CH_3OH



(d)

The **chemical formula** for **nickel (II) bromide** is

- A Ni_2Br
- B NiBr_2
- C N_2Br
- D NBr_2

(b)