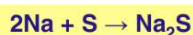




Name _____ Class _____ Date _____

1

Given the balanced equation:

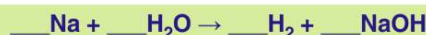


What is the **total number of moles of S** that reacted when **4.0 moles of Na** were completely consumed?

- A** 1.0 mole **C** 0.5 mole
B 2.0 moles **D** 4.0 moles

2

Given the unbalanced equation:



When the equation is correctly balanced using the *smallest* whole-number coefficients, the **coefficient for H₂O** is

- A** 1 **C** 3
B 2 **D** 4

3

When an equation is **correctly balanced**, it must show **conservation** of

- A** charge but not of mass
B mass but not of charge
C both charge and mass
D neither charge nor mass



4

Which **equation** is **correctly balanced**?

- A** $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
B $\text{Ca} + \text{Cl}_2 \rightarrow \text{CaCl}$
C $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
D $\text{Ca} + \text{Cl}_2 \rightarrow \text{Ca}_2\text{Cl}$

5



PREVIEW

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7

- A** addition
B hydrogenation
C polymerization
D substitution

- B** coefficients
C molecules
D moles of molecules

9

Given the **unbalanced** equation:



When the equation is balanced using the *smallest* whole-number coefficients, what is the **coefficient of Al**?

- A** 1 **C** 3
B 2 **D** 4

10

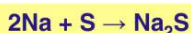
Which **equation** represents a **double replacement reaction**?

- A** $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$
B $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
C $\text{LiOH} + \text{HCl} \rightarrow \text{LiCl} + \text{H}_2\text{O}$
D $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$



ANSWER KEY

Given the balanced equation:

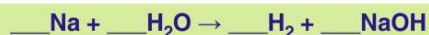


What is the **total number of moles of S** that reacted when **4.0 moles of Na** were completely consumed?

- A** 1.0 mole **C** 0.5 mole
B 2.0 moles **D** 4.0 moles

(b)

Given the unbalanced equation:



When the equation is correctly balanced using the *smallest* whole-number coefficients, the **coefficient for H₂O** is

- A** 1 **C** 3
B 2 **D** 4

(b)

When an equation is **correctly balanced**, it must show **conservation** of

- A** charge but not of mass
B mass but not of charge
C both charge and mass
D neither charge nor mass



(c)

Which **equation** is **correctly balanced**?

- A** $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
B $\text{Ca} + \text{Cl}_2 \rightarrow \text{CaCl}$
C $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
D $\text{Ca} + \text{Cl}_2 \rightarrow \text{Ca}_2\text{Cl}$

(c)



PREVIEW

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- A** reduction
B hydrogenation
C polymerization
D substitution

- C** molecules
D moles of molecules

Given the **unbalanced** equation:



When the equation is balanced using the *smallest* whole-number coefficients, what is the **coefficient of Al**?

- A** 1 **C** 3
B 2 **D** 4

(b)

Which **equation** represents a **double replacement reaction**?

- A** $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$
B $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
C $\text{LiOH} + \text{HCl} \rightarrow \text{LiCl} + \text{H}_2\text{O}$
D $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$

(c)