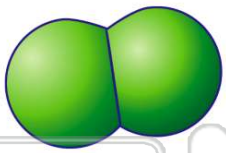
Elements - Set II

CHEM

Name _____ Class _____ Date _____

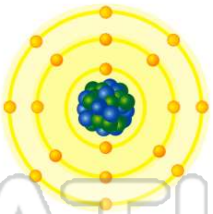
1 As two chlorine atoms **combine** to form a **molecule**, energy is

A absorbed
B released
C created
D destroyed



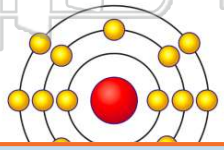
2 Which particle has the **same electron configuration** as a **potassium ion**?

A fluoride ion
B sodium ion
C neon atom
D argon atom



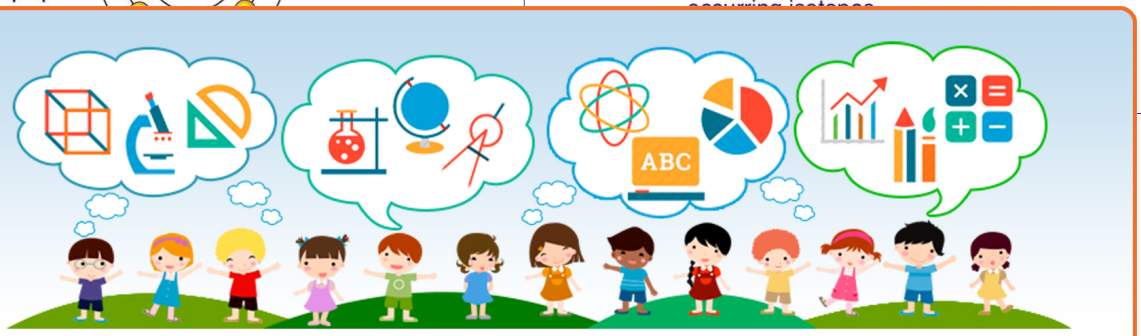
3 Which is an **electron configuration** for an atom of **chlorine** in the **excited state**?

A 2–8–7
B 2–8–8
C 2–8–6–1
D 2–8–7–1



4 The **atomic mass** of an **element** is calculated using the

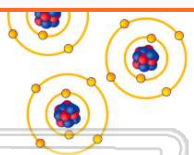
A atomic number and the ratios of its naturally occurring isotopes
B atomic number and the half-lives of each of its isotopes
C masses and the ratios of its naturally



5

7

B oxygen
C fluorine
D carbon



A decomposition
B transmutation
C substitution
D reduction

9 Which list of elements is arranged in order of **increasing** atomic radii?

A Li, Be, B, C
B Sr, Ca, Mg, Be
C Sc, Ti, V, Cr
D F, Cl, Br, I

10 Which isotope is most commonly used in the **radioactive dating** of the remains of organic materials?

A ^{14}C
B ^{16}N
C ^{32}P
D ^{37}K

Elements - Set II

CHEM

ANSWER KEY

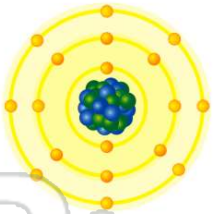
As two chlorine atoms **combine** to form a **molecule**, energy is

A absorbed
B released
C created
D destroyed

b

Which particle has the **same electron configuration** as a **potassium ion**?

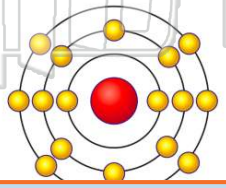
A fluoride ion
B sodium ion
C neon atom
D argon atom



d

Which is an **electron configuration** for an atom of **chlorine** in the **excited state**?

A 2–8–7
B 2–8–8
C 2–8–6–1
D 2–8–7–1

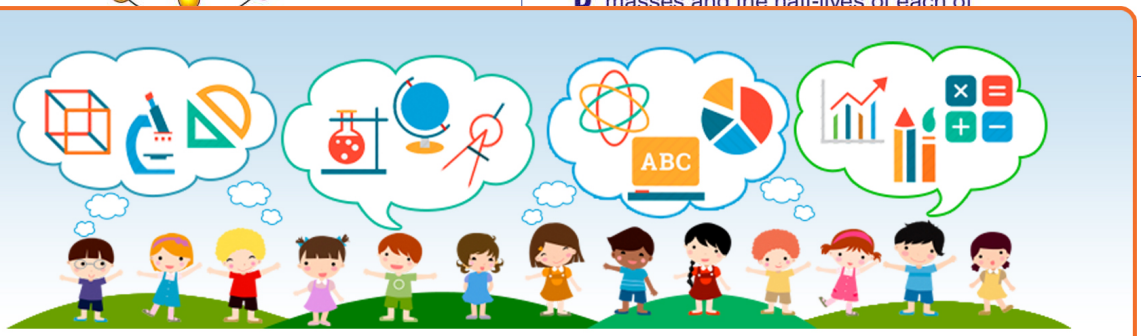


c

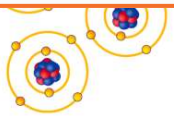
The **atomic mass** of an **element** is calculated using the

A atomic number and the ratios of its naturally occurring isotopes
B atomic number and the half-lives of each of its isotopes
C masses and the ratios of its naturally occurring isotopes
D masses and the half-lives of each of

c



C fluorine
D carbon



B transmutation
C substitution
D reduction

Which list of elements is arranged in order of **increasing** atomic radii?

A Li, Be, B, C
B Sr, Ca, Mg, Be
C Sc, Ti, V, Cr
D F, Cl, Br, I

d

Which isotope is most commonly used in the **radioactive dating** of the remains of organic materials?

A ^{14}C
B ^{16}N
C ^{32}P
D ^{37}K

a