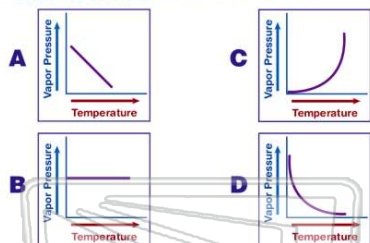




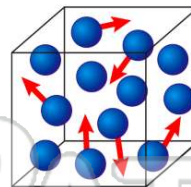
Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

1 Which graph best represents the **variation** in the **vapor pressure** of water as temperature changes?



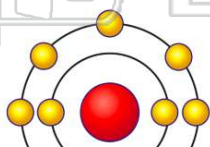
2 A 2.00-liter sample of a gas has a mass of 1.80 grams at STP. What is the **density**, in grams per liter, of this gas at STP?

- A 0.900
- B 1.80
- C 11.2
- D 22.4



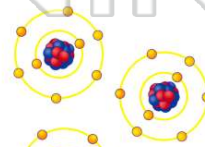
3 What is the **total number** of neon atoms contained in 20.2 grams of neon gas?

- A  $1.01 \times 10^{24}$
- B  $2.02 \times 10^{24}$
- C  $3.01 \times 10^{23}$
- D  $6.02 \times 10^{23}$



4 What is the **total number** of moles of **oxygen atoms** in 1 mole of  $N_2O_3$ ?

- A 1
- B 2
- C 3
- D 5



## PREVIEW

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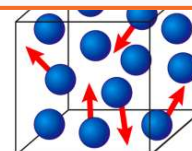
7

Sample of each of these gases

- A decrease
- B increase
- C remains the same

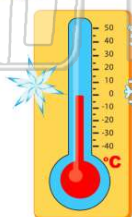
36	Ar	4
54	Xe	5
86	Rn	6

- B 0°C and 300 kPa
- C 150°C and 100 kPa
- D 150°C and 300 kPa



9 **Standard temperature** and a **pressure** of 0.5 atmosphere are equal to

- A 0°C and 380 torr
- B 32°C and 380 torr
- C 0°C and 760 torr
- D 32°C and 760 torr



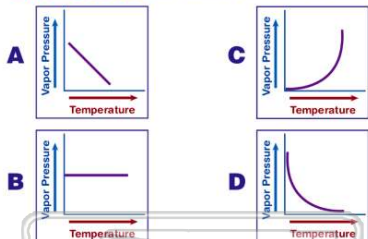
10 At STP,  $3 \times 10^{23}$  molecules of  $SO_2(g)$  occupy the **same** volume as

- A 1 mole of  $H_2(g)$
- B  $6 \times 10^{23}$  molecules of  $H_2(g)$
- C 0.5 mole of  $H_2(g)$
- D 4 grams of  $H_2(g)$



## ANSWER KEY

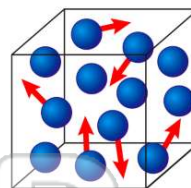
Which graph best represents the **variation** in the **vapor pressure** of water as temperature changes?



C

A 2.00-liter sample of a gas has a mass of 1.80 grams at STP. What is the **density**, in grams per liter, of this gas at STP?

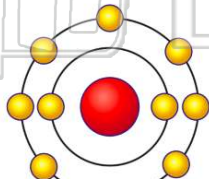
- A 0.900  
B 1.80  
C 11.2  
D 22.4



a

What is the **total number** of **neon atoms** contained in 20.2 grams of neon gas?

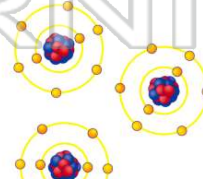
- A  $1.01 \times 10^{24}$   
B  $2.02 \times 10^{24}$   
C  $3.01 \times 10^{23}$   
D  $6.02 \times 10^{23}$



d

What is the **total number** of moles of **oxygen atoms** in 1 mole of  $\text{N}_2\text{O}_3$ ?

- A 1  
B 2  
C 3  
D 5



c



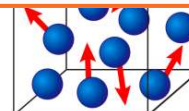
## PREVIEW

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- A decrease  
B increase  
C remains the same

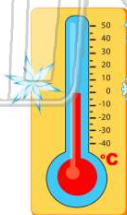
54	Xe	5
86	Rn	6

- C  $150^\circ\text{C}$  and 100 kPa  
D  $150^\circ\text{C}$  and 300 kPa



**Standard temperature** and a **pressure** of 0.5 atmosphere are equal to

- A  $0^\circ\text{C}$  and 380 torr  
B  $32^\circ\text{C}$  and 380 torr  
C  $0^\circ\text{C}$  and 760 torr  
D  $32^\circ\text{C}$  and 760 torr



a

At STP,  $3 \times 10^{23}$  **molecules** of  $\text{SO}_2(\text{g})$  occupy the **same volume** as

- A 1 mole of  $\text{H}_2(\text{g})$   
B  $6 \times 10^{23}$  molecules of  $\text{H}_2(\text{g})$   
C 0.5 mole of  $\text{H}_2(\text{g})$   
D 4 grams of  $\text{H}_2(\text{g})$

c