



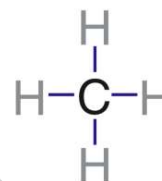
Name _____ Class _____ Date _____

1 What is the **gram molecular mass** of a gas that has a **density** of 5.00 grams per liter at STP?

- A 27.4 g
- B 56.0 g
- C 112 g
- D 223 g

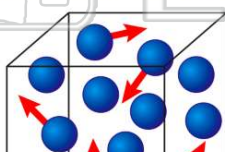
2 **Natural gas** is mostly comprised of

- A butane
- B ethane
- C methane
- D propane



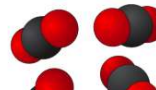
3 The **temperature** of a sample of nitrogen gas is a measure of the molecules' average

- A activation energy
- B potential energy
- C kinetic energy
- D ionization energy



4 The pressure on a 200-milliliter sample of $\text{CO}_2(\text{g})$ at constant temperature is **increased** from 600 torr to 1,200 torr. What is the **new volume** of the gas?

- A 100 mL
- B 300 mL
- C 400 mL
- D 600 mL

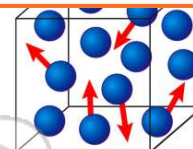


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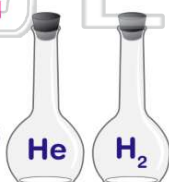
- 5
- A 200 torr
 - B 400 torr
 - C 2000 torr
 - D 4000 torr

- 6
- A 44.0 g
 - B 64.0 g
 - C 80.0 g
 - D 100.0 g



9 A sealed flask containing 1.0 mole of $\text{H}_2(\text{g})$ and a sealed flask containing 2.0 moles of $\text{He}(\text{g})$ are at the **same** temperature. **The two gases must have equal**

- A masses
- B volumes
- C average kinetic energies
- D numbers of molecules



10 Two **basic properties** of the **gas phase** are

- A a definite shape and a definite volume
- B a definite shape but no definite volume
- C no definite shape but a definite volume
- D no definite shape and no definite volume



ANSWER KEY

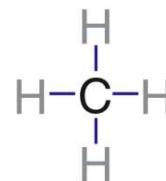
What is the **gram molecular mass** of a gas that has a **density** of **5.00 grams per liter** at STP?

- A 27.4 g
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- C 112 g
- D 223 g

C

Natural gas is mostly comprised of

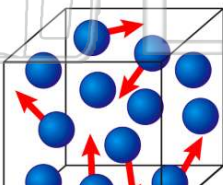
- A butane
- B ethane
- C methane
- D propane



C

The **temperature** of a sample of nitrogen gas is a measure of the molecules' average

- A activation energy
- B potential energy
- C kinetic energy
- D ionization energy



C

The pressure on a 200-milliliter sample of $\text{CO}_2(\text{g})$ at constant temperature is **increased** from 600 torr to 1,200 torr.

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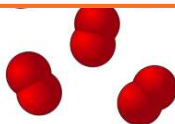
a



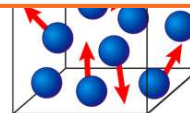
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- A 200 torr
- B 400 torr
- C 2000 torr
- D 4000 torr

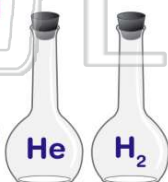


- C 64.0 g
- D 80.0 g



A sealed flask containing **1.0 mole** of $\text{H}_2(\text{g})$ and a sealed flask containing **2.0 moles** of $\text{He}(\text{g})$ are at the **same** temperature. **The two gases must have equal**

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C

Two **basic properties** of the **gas phase** are

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d