



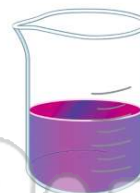
Name _____ Class _____ Date _____

1 In an **oxidation-reduction reaction**, **reduction** is defined as the

- A loss of protons
- B gain of protons
- C loss of electrons
- D gain of electrons

2 What is the **oxidation number** assigned to **manganese** in KMnO_4 ?

- A +7
- B +2
- C +3
- D +4



3 Which of the following **aqueous solutions** is the best **conductor of electricity**?

- A 0.10 M CH_3OH
- B 1.0 M CH_3OH
- C 0.10 M NaOH
- D 1.0 M NaOH



4 A **chemical cell** is made up of two half-cells connected by an external conductor and a salt bridge. The **function of the salt bridge** is to

- A permit the migration of ions
- B permit the mixing of solutions
- C prevent the migration of ions



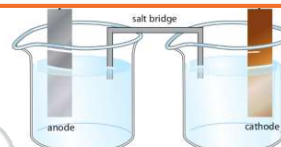
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7 In an **oxidation-reduction reaction**, the **oxidation number** of the **oxidizing agent**

- A decreases
- B increases
- C remains the same

- B increases
- C remains the same

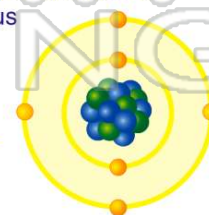


9 In an **oxidation-reduction reaction**, the **oxidation number** of the **oxidizing agent**

- A decreases
- B increases
- C remains the same

10 As ^{14}C decays to ^{14}N , the number of **protons** in the nucleus

- A decreases
- B increases
- C remains the same





ANSWER KEY

In an **oxidation-reduction reaction**, **reduction** is defined as the

- A loss of protons
- B gain of protons
- C loss of electrons
- D gain of electrons

(d)

What is the **oxidation number** assigned to **manganese** in KMnO_4 ?

- A +7
- B +2
- C +3
- D +4



(a)

Which of the following **aqueous solutions** is the best **conductor of electricity**?

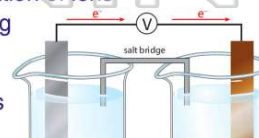
- A 0.10 M CH_3OH
- B 1.0 M CH_3OH
- C 0.10 M NaOH
- D 1.0 M NaOH



(d)

A **chemical cell** is made up of two half-cells connected by an external conductor and a salt bridge. The **function of the salt bridge** is to

- A permit the migration of ions
- B permit the mixing of solutions
- C prevent the migration of ions
- D prevent the flow

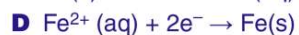


(a)



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C remains the same



In an **oxidation-reduction reaction**, the **oxidation number** of the **oxidizing agent**

- A decreases
- B increases
- C remains the same

(a)

As ^{14}C decays to ^{14}N , the number of **protons** in the nucleus

- A decreases
- B increases
- C remains the same

(b)

