



Name _____ Class _____ Date _____

1 What happens when **NaCl(s)** is **dissolved** in water?

- A Cl⁻ ions are attracted to the oxygen atoms of water molecules.
- B Na⁺ ions are attracted to the oxygen atoms of water molecules.
- C Cl⁻ ions are repelled by the hydrogen atoms of water molecules.
- D Na⁺ ions are repelled by the oxygen atoms of water molecules.

2 If **equal volumes** of **0.1 M NaOH** and **0.1 M HCl** are **mixed**, the resulting solution will contain a salt and

- A HCl
- B NaOH
- C H₂O
- D NaCl



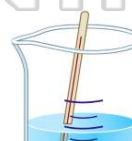
3 The **[H₃O⁺]** of a solution is **1 x 10⁻⁸**. This solution has a **pH** of

- A 6, which is acidic
- B 8, which is basic
- C 6, which is basic
- D 8, which is acidic



4 Which compound is **least soluble** in 100 grams of water at 40°C?

- A SO₂
- B NaCl
- C KClO₃
- D NH₄Cl



PREVIEW

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- B homogeneous compound
- C heterogeneous mixture
- D homogeneous mixture

- A organic
- B inorganic
- C basic
- D acidic

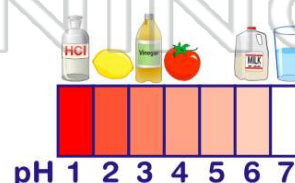


9 What is the **H₃O⁺ ion concentration** of a solution that has an OH⁻ ion concentration of 1.0 x 10⁻³ M?

- A 1.0 x 10⁻³ M
- B 1.0 x 10⁻⁷ M
- C 1.0 x 10⁻¹¹ M
- D 1.0 x 10⁻¹⁴ M

10 What is the **pH** of a **0.0001 M** aqueous solution of **HCl**?

- A 1
- B 2
- C 3
- D 4





ANSWER KEY

What happens when **NaCl(s)** is **dissolved** in water?

- A Cl⁻ ions are attracted to the oxygen atoms of water molecules.
- B Na⁺ ions are attracted to the oxygen atoms of water molecules.
- C Cl⁻ ions are repelled by the hydrogen atoms of water molecules.
- D Na⁺ ions are repelled by the oxygen atoms of water molecules.

(b)

If **equal volumes** of **0.1 M NaOH** and **0.1 M HCl** are **mixed**, the resulting solution will contain a salt and

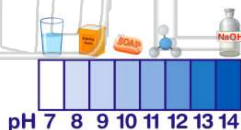
- A HCl
- B NaOH
- C H₂O
- D NaCl



(c)

The **[H₃O⁺]** of a solution is **1 x 10⁻⁸**. This solution has a **pH** of

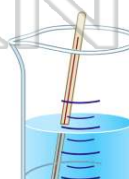
- A 6, which is acidic
- B 8, which is basic
- C 6, which is basic
- D 8, which is acidic



(b)

Which compound is **least soluble** in 100 grams of water at 40°C?

- A SO₂
- B NaCl
- C KClO₃
- D NH₄Cl



(a)



PREVIEW

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D homogeneous mixture

- B inorganic
- C basic
- D acidic



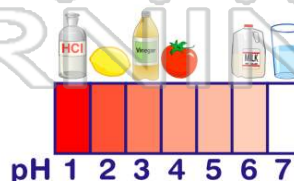
What is the **H₃O⁺ ion concentration** of a solution that has an OH⁻ ion concentration of 1.0 x 10⁻³ M?

- A 1.0 x 10⁻³ M
- B 1.0 x 10⁻⁷ M
- C 1.0 x 10⁻¹¹ M
- D 1.0 x 10⁻¹⁴ M

(c)

What is the **pH** of a **0.0001 M** aqueous solution of **HCl**?

- A 1
- B 2
- C 3
- D 4



(d)