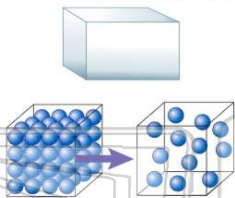




Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

1 Under the same conditions of temperature and pressure, which sample contains particles having the **lowest entropy**?

- A  $\text{CO}_2(\text{g})$
- B  $\text{CO}_2(\text{l})$
- C  $\text{CO}_2(\text{s})$
- D  $\text{CO}_2(\text{aq})$

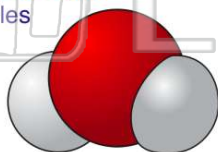


2 At 25°C, **iodine** exists as a **solid** and **bromine** as a **liquid**. This difference in phase is due to the presence of stronger

- A covalent bonds in iodine
- B covalent bonds in bromine
- C van der Waals forces in iodine
- D van der Waals forces in bromine

3 The primary forces of **attraction** between water molecules in  $\text{H}_2\text{O}(\text{l})$  are

- A ionic bonds
- B hydrogen bonds
- C molecule-ion attractions



4 Which type of **solid** is silicon carbide ( $\text{SiC}$ )?

- A ionic
- B molecular
- C network
- D metallic



## PREVIEW

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- 7
- A 1 g
  - B 10 g
  - C 100 g
  - D 1000 g

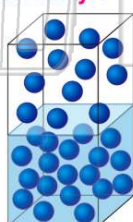


- B boiling point
- C heat of fusion
- D heat of vaporization



9 As the pressure on a gas confined above a liquid increases, the **solubility** of the gas in the liquid

- A decreases
- B increases
- C remains the same



10 As the temperature of a gas increases at constant pressure, the **volume** of the gas

- A decreases
- B increases
- C remains the same

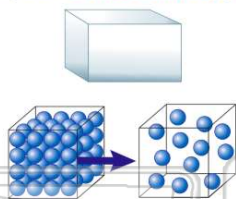




## ANSWER KEY

Under the same conditions of temperature and pressure, which sample contains particles having the **lowest entropy**?

- A  $\text{CO}_2(\text{g})$
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(C)

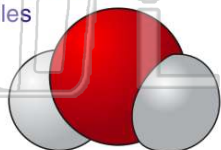
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(C)

The primary forces of **attraction** between water molecules in  $\text{H}_2\text{O}(\text{l})$  are

- A ionic bonds
- B hydrogen bonds
- C molecule-ion attractions
- D van der Waals forces



(B)

Which type of **solid** is silicon carbide ( $\text{SiC}$ )?

- A ionic
- B molecular
- C network
- D metallic

(C)



## PREVIEW

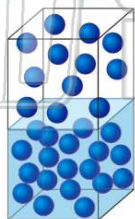
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- B 10 g
- C 100 g
- D 1000 g



As the pressure on a gas confined above a liquid increases, the **solubility** of the gas in the liquid

- A decreases
- B increases
- C remains the same



(B)

- C heat of fusion
- D heat of vaporization



As the temperature of a gas increases at constant pressure, the **volume** of the gas

- A decreases
- B increases
- C remains the same



(B)