



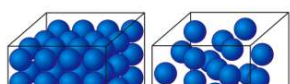
Name _____ Class _____ Date _____

1 If 4.00 moles of oxygen gas, 3.00 moles of hydrogen gas, and 1.00 mole of nitrogen gas are combined in a closed container at standard pressure, what is the partial pressure exerted by the hydrogen gas?

- A 1.00 atm C 3.00 atm
B 0.125 atm D 0.375 atm

3 The solid and liquid phases of water can exist in a state of equilibrium at 1 atmosphere of pressure and a temperature of

- A 0°C
B 100°C



2 Which statement correctly describes a sample of gas confined in a sealed container?

- A It always has a definite volume, and it takes the shape of the container.
B It takes the shape and the volume of any container in which it is confined.
C It has a crystalline structure.
D It consists of particles arranged in a regular geometric pattern.

4 As the pressure on the surface of a liquid decreases, the temperature at which the liquid will boil

- A decreases
B increases
C remains the same



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B increases
C remains the same



- B condensing of water vapor
C subliming of iodine
D melting of ice



9 Which substance can not be decomposed by a chemical change?

- A Ne
B N₂O
C HF
D H₂O



10 A real gas behaves more like an ideal gas when the gas molecules are

- A close and have strong attractive forces between them
B close and have weak attractive forces between them
C far apart and have strong attractive forces between them
D far apart and have weak attractive forces between them



ANSWER KEY

If 4.00 moles of oxygen gas, 3.00 moles of hydrogen gas, and 1.00 mole of nitrogen gas are combined in a closed container at standard pressure, what is the partial pressure exerted by the hydrogen gas?

- A 1.00 atm
- B 0.125 atm
- C 3.00 atm
- D 0.375 atm

(d)

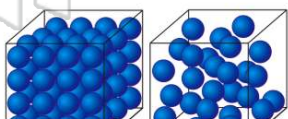
Which statement correctly describes a sample of gas confined in a sealed container?

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- B It takes the shape and the volume of any container in which it is confined.
- C It has a crystalline structure.
- D It consists of particles arranged in a regular geometric pattern.

(b)

The solid and liquid phases of water can exist in a state of equilibrium at 1 atmosphere of pressure and a temperature of

- A 0°C
- B 100°C
- C 273°C



(a)

As the pressure on the surface of a liquid decreases, the temperature at which the liquid will boil

- A decreases
- B increases
- C remains the same



(a)



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C remains the same



D melting of ice



Which substance can not be decomposed by a chemical change?

- A Ne
- B N₂O
- C HF
- D H₂O



(a)

A real gas behaves more like an ideal gas when the gas molecules are

- A close and have strong attractive forces
- B close and have weak attractive forces between them
- C far apart and have strong attractive forces between them
- D far apart and have weak attractive forces between them

(d)