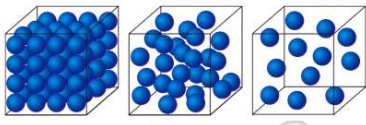




Name _____ Class _____ Date _____

- 1 Which phase has the **highest specific heat**?
- A solid
B liquid
C gas
- 

- 3 As the **pressure** of a fixed mass of gas is **increased** at constant temperature, the **density** of that gas
- A decreases
B increases
C remains the same

- 4 After a hot object is placed in an insulated container with a cold object, the hot object changes temperature and the cold object changes phase. The **total amount of internal energy** in the system will
- A decrease
B increase
- 



PREVIEW

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- B higher average internal kinetic energy
C greater mass
D greater specific heat

- A 20°C
B 22°C
C 40°C
D 60°C
- 

- 9 Equal masses of **copper**, **iron**, **lead**, and **silver** are heated from 20°C to 100°C. Which substance absorbs the **least amount of heat**?
- A lead
B iron
C copper
D silver

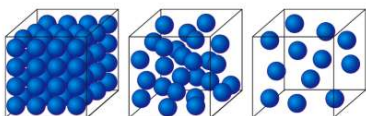
- 10 **Block A**, at **100°C**, and **block B**, at **50°C**, are brought together in a well-insulated container. The **internal energy of block A** will
- A decrease and the internal energy of block B will decrease
B decrease and the internal energy of block B will increase
C increase and the internal energy of block B will decrease
D increase and the internal energy of block B will increase



ANSWER KEY

Which phase has the **highest specific heat**?

- A solid
- B liquid
- C gas



(b)

In an **ideal gas**, **entropy** is a measure of the

- A volume of the molecules
- B mass of the molecules
- C forces of attraction between the molecules
- D disorder of the molecules

(d)

As the **pressure** of a fixed mass of gas is **increased** at constant temperature, the **density** of that gas

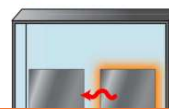
- A decreases
- B increases
- C remains the same

(b)

After a hot object is placed in an insulated container with a cold object, the hot object changes temperature and the cold object changes phase. The **total amount of internal energy** in the system will

- A decrease
- B increase
- C remain the same

(c)



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- C greater mass
- D greater specific heat

- B 22°C
- C 40°C
- D 60°C



Equal masses of **copper**, **iron**, **lead**, and **silver** are heated from 20°C to 100°C. Which substance absorbs the **least amount of heat**?

- A lead
- B iron
- C copper
- D silver

(a)

Block A, at **100°C**, and **block B**, at **50°C**, are brought together in a well-insulated container. The **internal energy of block A** will

- A decrease and the internal energy of block B will decrease
- B decrease and the internal energy of block B will increase
- C increase and the internal energy of block B will decrease
- D increase and the internal energy of block B will increase

(b)