



Name _____ Class _____ Date _____

1 What has **taken place** in this picture?

- A a chemical reaction
- B a physical change
- C an exothermic reaction
- D a chemical change



2 In order for a **chemical change** to take place, _____.

- A products must turn into reactants
- B reactants must turn into products
- C heat must be given off
- D heat must be added



3 What kind of **reaction** has occurred in this picture?

- A synthesis
- B decomposition
- C single replacement



4 In a chemical equation, the **reactants** are on the **left** side.
True or false?

- A true $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
- B false

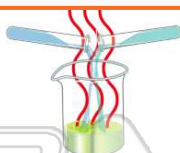


PREVIEW

Please [Sign In](#) or [Sign Up](#) to download the printable version of this worksheet

7

A endothermic reaction
B physical reaction
C physical change
D exothermic reaction



- A The reaction is going from the right to the left.
- B The reactants are on the right side.
- C The products are different than the reactants.
- D The products are on the left side.

9 Which statement is correct about the **chemical reaction** shown below?



- A The mass of the reactants equals the mass of the products.
- B The mass of the reactants is less than the mass of the products.
- C The mass of the products is less than the mass of the reactants.
- D Mass has changed into energy.

10 What is the explanation for an **acid** and a **base** changing into **salt** and **water**?

- A A physical change has taken place. $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- B A synthesis reaction has taken place.
- C A decomposition reaction has taken place.
- D A chemical change has taken place.



ANSWER KEY

What has **taken place** in this picture?

- A a chemical reaction
- B a physical change
- C an exothermic reaction
- D a chemical change



(b)

In order for a **chemical change** to take place, _____.

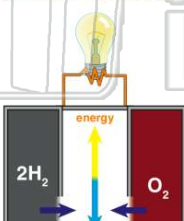
- A products must turn into reactants
- B reactants must turn into products
- C heat must be given off
- D heat must be added



(b)

What kind of **reaction** has occurred in this picture?

- A synthesis
- B decomposition
- C single replacement
- D double



(a)

In a chemical equation, the **reactants** are on the **left** side.

True or false?

- A true $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
- B false

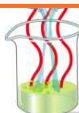
(a)



PREVIEW

Please [Sign In](#) or [Sign Up](#) to download the printable version of this worksheet

- B physical reaction
- C physical change
- D exothermic reaction



- B The reactants are on the right side.
- C The products are different than the reactants.
- D The products are on the left side.

Which statement is correct about the **chemical reaction** shown below?



- A The mass of the reactants equals the mass of the products.
- B The mass of the reactants is less than the mass of the products.
- C The mass of the products is less than the mass of the reactants.
- D Mass has changed into energy.

(a)

What is the explanation for an **acid** and a **base** changing into **salt** and **water**?

- A A physical change has taken place.
- B A synthesis reaction has taken place.
- C A decomposition reaction has taken place.
- D A chemical change has taken place.



(d)