



Name _____ Class _____ Date _____

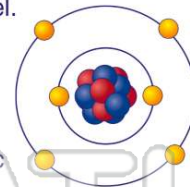
1 In this modern drawing of an atom, what is present that would not be drawn in **Bohr's model** of the atom?

- A electrons
- B protons
- C neutrons
- D orbits



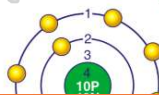
2 In 1913, the model of the atom was revised to show **electrons circling** the nucleus in **specific orbits**. This is called _____ model.

- A Bohr's
- B Rutherford's
- C Dalton's
- D the modern atomic



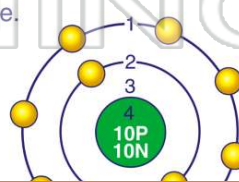
3 **Protons** are positively charged, **electrons** are negatively charged, and **neutrons** are neutral. Using the diagram below, determine the **overall charge**.

- A positive
- B negative



4 Point #4 in the diagram has _____ charge.

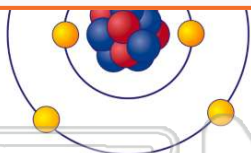
- A no
- B an electron
- C a negative
- D a positive



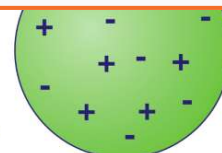
PREVIEW

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7 A _____
a _____
A _____
B compound
C molecule
D chemical bond

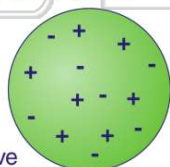


- A the modern
- B the Rutherford
- C the Bohr
- D none of the above



9 The diagram below of an **earlier model** of an atom shows **protons** and **electrons** existing together in the center of an atom. If this model was correct, then the atom shown would have which charge?

- A positive
- B negative
- C none
- D both positive and negative



10 To convert one atom to an **isotope** of the **same element**, the number of _____ would have to be changed.

- A electrons
- B neutrons
- C protons
- D elements



ANSWER KEY

In this modern drawing of an atom, what is present that would not be drawn in **Bohr's model** of the atom?

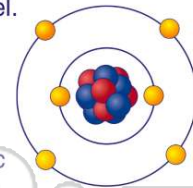
- A electrons
- B protons
- C neutrons
- D orbits



(C)

In 1913, the model of the atom was revised to show **electrons circling** the nucleus in **specific orbits**. This is called _____ model.

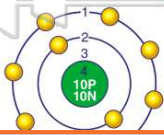
- A Bohr's
- B Rutherford's
- C Dalton's
- D the modern atomic



(a)

Protons are positively charged, **electrons** are negatively charged, and **neutrons** are neutral. Using the diagram below, determine the **overall charge**.

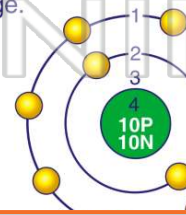
- A positive
- B negative
- C neutral
- D



(C)

Point #4 in the diagram has _____ charge.

- A no
- B an electron
- C a negative
- D a positive



(d)



PREVIEW

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- D chemical bond

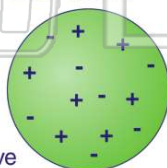


- B the Rutherford
- C the Bohr
- D none of the above



The diagram below of an **earlier model** of an atom shows **protons** and **electrons** existing together in the center of an atom. If this model was correct, than the atom shown would have which charge?

- A positive
- B negative
- C none
- D both positive and negative



(C)

To convert one atom to an **isotope** of the **same element**, the number of _____ would have to be changed.

- A electrons
- B neutrons
- C protons
- D elements

(b)