



Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

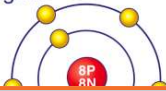
1 According to the **modern theory** of atomic structure, most of an atom is made up of \_\_\_\_\_.

- A protons
- B neutrons
- C electrons
- D empty space



3 In the diagram below, the **electron orbits** are sometimes referred to as **energy levels**. What is the explanation for this?

- A Electrons are positively charged.
- B Electrons have no charge.
- C Electrons need energy to stay near the nucleus.

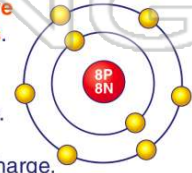


2 Two atoms of the **same element** must have the same \_\_\_\_\_.

- A number of protons
- B mass number
- C number of neutrons
- D size

4 Using the diagram below, explain why the **electrons do not totally leave** the area around the **nucleus**.

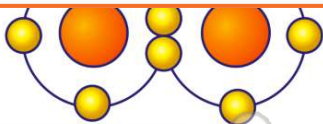
- A Positive and negative particles repel each other.
- B The electrons and the nucleus have the same charge.



## PREVIEW

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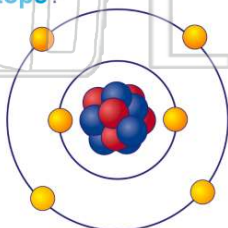
7  
A  
B neutrons  
C electrons  
D nucleus



- A the number of neutrons it has
- B the number of protons it has
- C the shape of the atom
- D the energy of its electrons

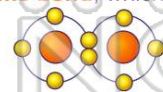
9 Which part of atom is involved in the **formation of an isotope**?

- A its protons
- B its electrons
- C its neutrons
- D its electron orbits



10 When two or more **atoms bond**, which of these is correct?

- A Bonded atoms of different elements make up compounds.
- B Bonded compounds of different elements make up atoms.
- C Elements bond to make up atoms.
- D Compounds are made from only one element.





## ANSWER KEY

According to the **modern theory** of atomic structure, most of an atom is made up of \_\_\_\_\_.

- A protons
- B neutrons
- C electrons
- D empty space



(d)

Two atoms of the **same element** must have the same \_\_\_\_\_.

- A number of protons
- B mass number
- C number of neutrons
- D size

(a)

In the diagram below, the **electron orbits** are sometimes referred to as **energy levels**. What is the explanation for this?

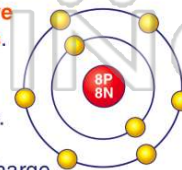
- A Electrons are positively charged.
- B Electrons have no charge.
- C Electrons need energy to stay near the nucleus.
- D Electrons need energy to stay away from the nucleus.



(d)

Using the diagram below, explain why the **electrons do not totally leave** the area around the **nucleus**.

- A Positive and negative particles repel each other.
- B The electrons and the nucleus have the same charge.
- C Neutrons pull the electrons in.
- D Positive and negative particles attract.



(d)



## PREVIEW

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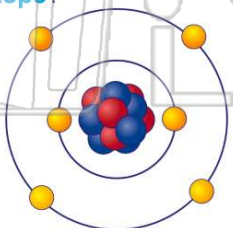
D nucleus



- C the shape of the atom
- D the energy of its electrons

Which part of atom is involved in the **formation of an isotope**?

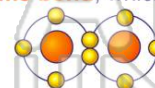
- A its protons
- B its electrons
- C its neutrons
- D its electron orbits



(c)

When two or more **atoms bond**, which of these is correct?

- A Bonded atoms of different elements make up compounds.
- B Bonded compounds of different elements make up atoms.
- C Elements bond to make up atoms.
- D Compounds are made from only one element.



(a)