



Name _____ Class _____ Date _____

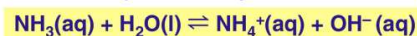
1

Which **acid-base pair** will always undergo a reaction that produces a **neutral solution**?

- A a weak acid and a weak base
- B a weak acid and a strong base
- C a strong acid and a weak base
- D a strong acid and a strong base

2

Given the **equilibrium** system:



According to the Brønsted-Lowry theory, the **H₂O(l)** acts as

- A a base, by receiving a proton
- B a base, by donating a proton
- C an acid, by receiving a proton
- D an acid, by donating a proton

3

How many milliliters of **0.010 M NaOH** are required to **exactly neutralize 20.0 milliliters** of **0.020 M HCl**?

4

Which of the following is the **weakest Brønsted-Lowry acid**?

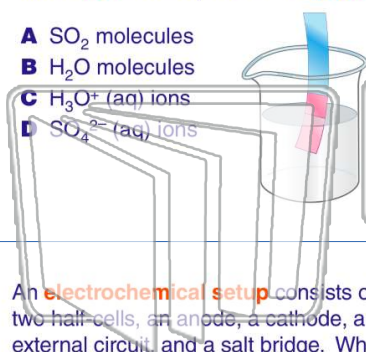
5

PREVIEW

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7

- A SO₂ molecules
- B H₂O molecules
- C H₃O⁺ (aq) ions
- D SO₄²⁻ (aq) ions

**9**

An **electrochemical setup** consists of two half-cells, an anode, a cathode, an external circuit, and a salt bridge. When a reaction occurs, **ion migration** takes place through the

- A anode
- B cathode
- C salt bridge
- D external circuit

and OH⁻ are mixed?

- A oxidation
- B reduction
- C hydrolysis
- D neutralization

10

According to the **Brønsted-Lowry theory**, an **acid** is any species that

- A releases hydroxide ions into solution
- B releases oxide ions into solution
- C donates protons to another species
- D accepts protons from another species

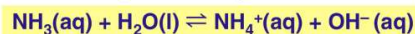


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A H₂SO₄

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