

## **Chemical Equations**



Name	Cla	ass	Date
	Given the unbalanced equation:  Al(s) + $O_2(g) \rightarrow Al_2O_3(s)$ When this equation is correctly palances using smallest whole numbers, what is the cosmological $O_2(g)$ ?	Which chemical ecorrectly balanced  A H <sub>2</sub> (g) + O <sub>2</sub> (g) -  B N <sub>2</sub> (g) + H <sub>2</sub> (g) -  C 2NaCl(s) - Na	d? H <sub>2</sub> O(g) NH <sub>3</sub> (g) (s) + C <sub>2</sub> (g)
3	Given the balanced equation:		NG
		ABC	
5	PREV		
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9	C <sub>2</sub> H <sub>2</sub> + 2Cl <sub>2</sub> $\rightarrow$ C <sub>2</sub> H <sub>2</sub> Cl <sub>4</sub> This reaction is best classified as  addition  begin{align*} Given the balanced equation:  C <sub>3</sub> H <sub>8</sub> (g) + 5O <sub>2</sub> (g) $\rightarrow$ 3CO <sub>2</sub> (g) + 4H <sub>2</sub> O(g)  What is the total number of liters of CO <sub>2</sub> (g) produced when 20.0 liters of O <sub>2</sub> (g) are	What is the total number of the total number o	umber of moles of when 3 moles of H2(g) umed?  moles moles  deguation:  moles moles of H2(g)  deguation:



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Name Class Date Given the unbalanced equation: Which chemical equation is correctly balanced?  $AI(s) + \underline{\hspace{1cm}} O_2(g) \rightarrow \underline{\hspace{1cm}} AI_2O_3(s)$ When this equation is correctly balance  $A H_2(g) + O_2(g) \rightarrow H_2O(g)$ using smalle  $B N_2(g) + H_2(g) \rightarrow NH_3(g)$ D the coefficient of Q<sub>o</sub>(g) C 2NaCl(s) Na(s) + Cl2(d) 2KCI(s) 6 Given the balanced equation: 3 Given the balanced equation:  $I_2(s)$  + energy  $\rightarrow I_2(g)$ 2C + 3H<sub>2</sub> → C<sub>2</sub>H<sub>6</sub> As a sample of I2(s) sublimes to I2(g), the entropy of the sample What is the total number of moles of C D 5  $(\mathbf{C})$ **PREVIEW** Please Sign In or Sign Up to download the printable version of this worksheet 7  $\Pi_2(g) + \text{Cl}_2(g) \rightarrow 2\Pi\text{Cl}(g)$ C<sub>2</sub>H<sub>2</sub> + 2Cl<sub>2</sub> → C<sub>2</sub>H<sub>2</sub>Cl<sub>4</sub> What is the total number of moles of This reaction is best classified as HCI(g) produced when 3 moles of H2(g) D A addition **B** esterification A 5 moles 3 moles fermentation B 2 moles D 6 moles substitution 9 Given the balanced equation: Siven the balanced equation  $C_3H_8(g) + 5O_2(g) \rightarrow 3CO_2(g) + 4H_2O(g)$  $Fe(s) + CuSO_4(aq) \rightarrow FeSO_4(aq) + Cu(s)$ What is the total number of liters of CO<sub>2</sub>(g) What total mass of iron is necessary produced when 20.0 liters of O2(g) are to produce 1.00 mole of copper? B completely consumed? A 26.0 g C 112 g C 3.00 L A 12.0 L **B** 55.8 q **D** 192 g B 22.4 L **D** 5.00 L