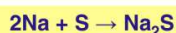




Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

1

Given the balanced equation:



What is the **total number of moles of S** that reacted when **4.0 moles of Na** were completely consumed?

- A 1.0 mole      C 0.5 mole  
B 2.0 moles    D 4.0 moles

2

Given the unbalanced equation:



When the equation is correctly balanced using the **smallest** whole-number coefficients, the **coefficient for H<sub>2</sub>O** is

- A 1      C 3  
B 2      D 4

3

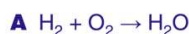
When an equation is **correctly balanced**, it must show **conservation of**

A charge but not of mass



4

Which **equation** is **correctly balanced**?



5



## PREVIEW

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7

Which **type of reaction** does this equation represent?

- A addition  
B hydrogenation  
C polymerization  
D substitution

the **same number** of

- A atoms  
B coefficients  
C molecules  
D moles of molecules

9

Given the **unbalanced** equation:



When the equation is balanced using the **smallest** whole-number coefficients, what is the **coefficient of Al**?

- A 1      C 3  
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10

Which **equation** represents a **double replacement reaction**?

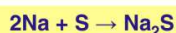
- A  $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$   
B  $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$   
C  $\text{LiOH} + \text{HCl} \rightarrow \text{LiCl} + \text{H}_2\text{O}$   
D  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$



Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

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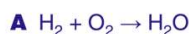
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