



Name _____ Class _____ Date _____

1

What is the **gram molecular mass** of a gas that has a **density of 5.00 grams per liter** at STP?

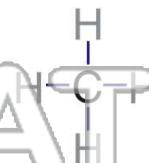
- A 27.4 g
- B 56.0 g
- C 112 g
- D 223 g



2

Natural gas is mostly comprised of

- A butane
- B ethane
- C methane
- D propane



NEW PATH LEARNING

3

The **temperature** of a sample of nitrogen gas is a measure of the molecules' average

A activation energy



4

The pressure on a 200-milliliter sample of $\text{CO}_2(\text{g})$ at constant temperature is **increased** from 600 torr to 1,200 torr.

What is the **new volume** of the gas?

5



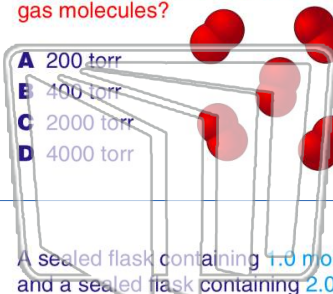
PREVIEW

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7

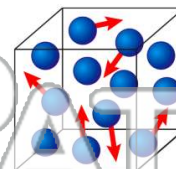
nitrogen gas and 2 moles of oxygen gas. What **pressure** is exerted by the **oxygen** gas molecules?

- A 200 torr
- B 400 torr
- C 2000 torr
- D 4000 torr



gram-molecular mass of this gas?

- A 28.0 g
- B 44.0 g
- C 64.0 g
- D 80.0 g



NEW PATH LEARNING

9

A sealed flask containing 1.0 mole of $\text{H}_2(\text{g})$ and a sealed flask containing 2.0 moles of $\text{He}(\text{g})$ are at the **same** temperature. The **two gases must have equal**

- A masses
- B volumes
- C average kinetic energies
- D numbers of molecules



10

Two **basic properties** of the **gas phase** are

- A a definite shape and a definite volume
- B a definite shape but no definite volume
- C no definite shape but a definite volume
- D no definite shape and no definite volume

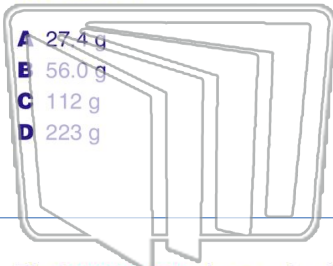


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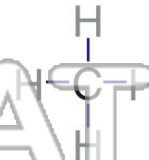
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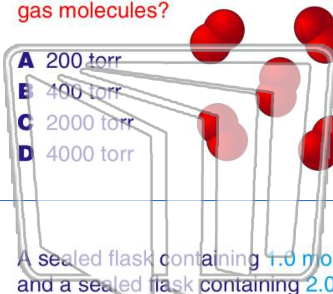
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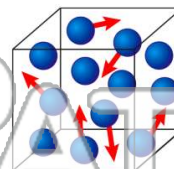
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