



Name _____ Class _____ Date _____

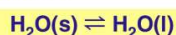
- 1 Which **equilibrium constant** indicates the **highest concentration** of product?

A $K_{eq} = 1 \times 10^{-1}$
 B $K_{eq} = 2 \times 10^{-2}$
 C $K_{eq} = 3 \times 10^{-3}$
 D $K_{eq} = 4 \times 10^{-4}$

- 2 Which statement best describes a **chemical reaction** when it reaches **equilibrium**?

- A The concentrations of reactants and products are the same.
 B The concentrations of the reactants decrease to zero.
 C The forward and reverse reaction rates are the same.
 D The forward reaction rate decreases to zero.

- 3 Given the equation:



At which temperature will equilibrium exist

- 4 Which **statement** correctly describes a chemical reaction at **equilibrium**?

- A The concentrations of the products and

5



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constant, the **activation energy** of the forward reaction

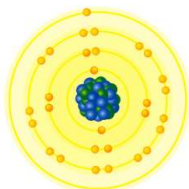
- A decreases
 B increases
 C remains the same

- A a chemical change, only
 B a physical change, only
 C both a chemical and a physical change
 D neither a chemical nor a physical change

9

An **increase** in the **average kinetic energy** of a sample of copper atoms occurs with an **increase** in

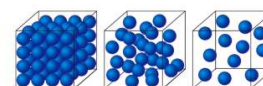
- A concentration
 B temperature
 C pressure
 D volume



10

Which **phase change** represents a **decrease in entropy**?

- A solid to liquid
 B gas to liquid
 C liquid to gas
 D solid to gas





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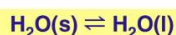
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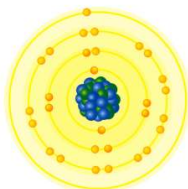
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