

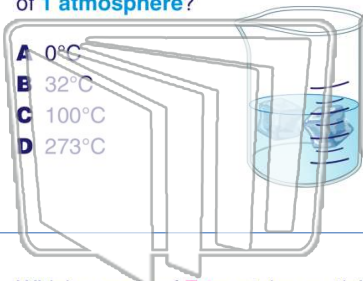


Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

**1**

What is the **equilibrium temperature** of an ice-water mixture at a pressure of **1 atmosphere**?

- A 0°C
- B 32°C
- C 100°C
- D 273°C

**2**

The list below shows four samples: **1, 2, 3, and 4**.

- (1) HCl(aq)      (3) HCl(g)
- (2) NaCl(aq)    (4) NaCl(s)

Which samples are **substances**?

- A 1 and 2      C 3 and 2
- B 1 and 3      D 3 and 4

**3**

Which sample of **Fe** contains particles having the **highest** average kinetic energy?

- A 5 g at 10°C

**4**

Which substance at **STP** **conducts electricity** because the substance contains **mobile electrons**?

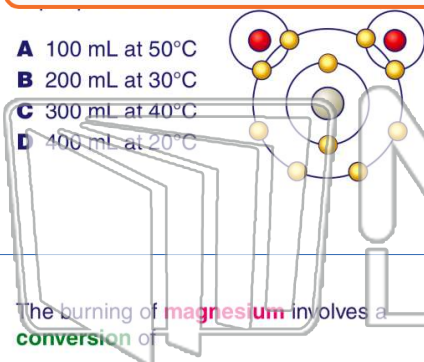
**5**

## PREVIEW

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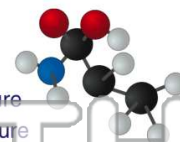
**7**

- A 100 mL at 50°C
- B 200 mL at 30°C
- C 300 mL at 40°C
- D 400 mL at 20°C



**chemically combined** in a definite ratio?

- A a solution
- B a compound
- C a homogeneous mixture
- D a heterogeneous mixture

**9**

The burning of **magnesium** involves a **conversion** of

- A chemical energy to mechanical energy
- B chemical energy to heat energy
- C heat energy to chemical energy
- D heat energy to mechanical energy

**10**

Which statement explains why **H<sub>2</sub>O** has a **higher boiling point** than **N<sub>2</sub>**?

- A H<sub>2</sub>O has greater molar mass than N<sub>2</sub>.
- B H<sub>2</sub>O has less molar mass than N<sub>2</sub>.
- C H<sub>2</sub>O has stronger intermolecular forces than N<sub>2</sub>.
- D H<sub>2</sub>O has weaker intermolecular forces than N<sub>2</sub>.

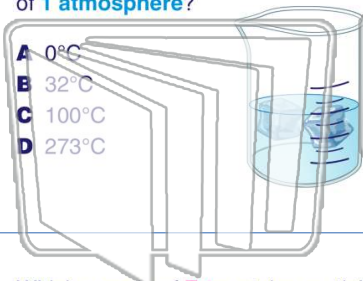


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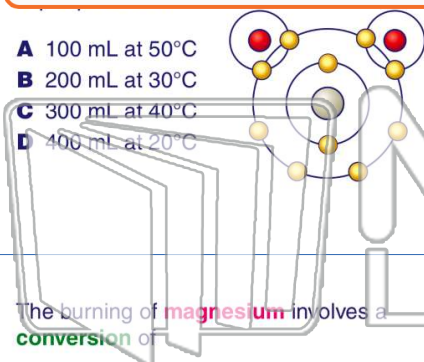
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