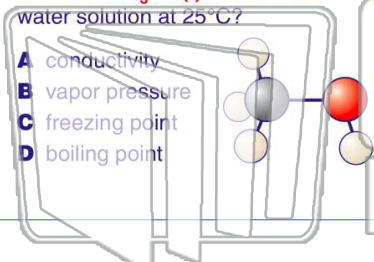




Name _____ Class _____ Date _____

1 Which property of a **distilled water solution** will **not** be affected by adding 50 mL of **CH₃OH(l)** to 100 mL of the water solution at 25°C?

- A conductivity
- B vapor pressure
- C freezing point
- D boiling point



2 What is the **gram formula mass** of **Ca₃(PO₄)₂**?

- A 196 g
- B 214 g
- C 245 g
- D 310 g



NEW PATH LEARNING

3 What is the **total number of grams of HI** in 0.500 liter of 1.00 M HI?

- A 1.00 g

4 Approximately **how many calories of heat** are needed to completely change 10 grams of **ice to water** at the melting point?

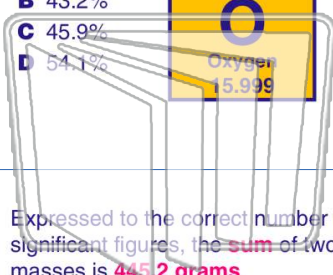
PREVIEW

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5

7

- A 21.6%
- B 43.2%
- C 45.9%
- D 54.1%



9 Expressed to the correct number of significant figures, the **sum** of two masses is **445.2 grams**.

Which two masses produce this answer?

- A 210.10 g + 235.100 g
- B 210.100 g + 235.10 g
- C 210.1 g + 235.1 g
- D 210.10 g + 235.10 g

NaOH?

- A 1.0 L
- B 2.0 L
- C 3.0 L
- D 4.0 L



10

A **hydrated salt** is a solid that includes water molecules within its crystal structure. A student heated a **9.10-gram** sample of a hydrated salt to a constant mass of **5.41 grams**.

What **percent** by mass of water did the **salt** contain?

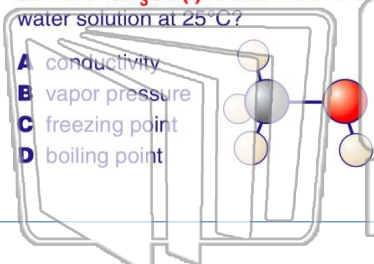
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Name _____ Class _____ Date _____

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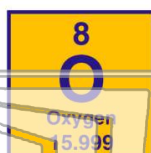


PREVIEW

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