

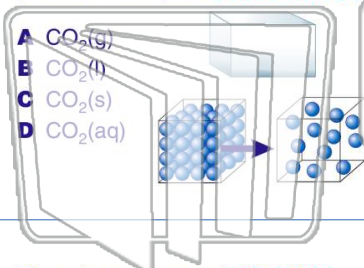


Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

1

Under the same conditions of temperature and pressure, which sample contains particles having the **lowest entropy**?

- A  $\text{CO}_2(\text{g})$
- B  $\text{CO}_2(\text{l})$
- C  $\text{CO}_2(\text{s})$
- D  $\text{CO}_2(\text{aq})$



2

At  $25^\circ\text{C}$ , **iodine** exists as a **solid** and **bromine** as a **liquid**. This difference in phase is due to the presence of stronger

- A covalent bonds in iodine
- B covalent bonds in bromine
- C van der Waals forces in iodine
- D van der Waals forces in bromine

3

The primary forces of **attraction** between water molecules in  $\text{H}_2\text{O}(\text{l})$  are

4

Which type of **solid** is silicon carbide ( $\text{SiC}$ )?

- A ionic

5



## PREVIEW

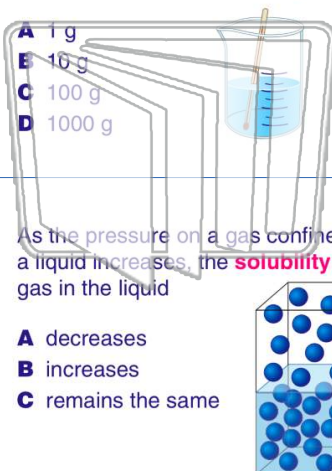
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7

water **absorbs** 100 calories of heat.

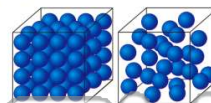
What is the mass of the sample?

- A 1 g
- B 10 g
- C 100 g
- D 1000 g



**equilibrium** is called its

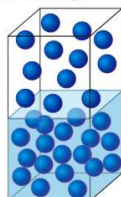
- A melting point
- B boiling point
- C heat of fusion
- D heat of vaporization



9

As the pressure on a gas confined above a liquid increases, the **solubility** of the gas in the liquid

- A decreases
- B increases
- C remains the same



10

As the temperature of a gas increases at constant pressure, the **volume** of the gas

- A decreases
- B increases
- C remains the same





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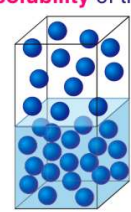
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