

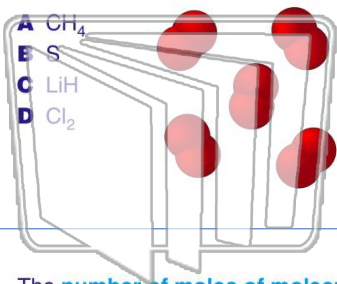


Name _____ Class _____ Date _____

1

One mole of O_2 has **approximately the same mass** as one mole of

- A CH_4
- B S
- C LiH
- D Cl_2



2

Which 1.0-mole sample at 1 atm has particles with the **greatest entropy**?

- A $CH_4(g)$ at $25^\circ C$
- B $H_2S(g)$ at $40^\circ C$
- C $CH_4(g)$ at 300 K
- D $H_2S(g)$ at 310 K

3

The **number of moles of molecules** in a **12.0-gram** sample of Cl_2 is

12.0

4

The percent by **mass** of **water** in the hydrate $Na_2SO_4 \cdot 10H_2O$ is closest to

A 18%

5

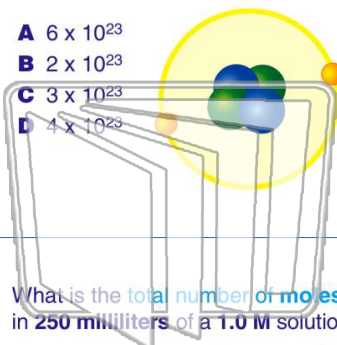


PREVIEW

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7

- A 6×10^{23}
- B 2×10^{23}
- C 3×10^{23}
- D 4×10^{23}



9

What is the **total number of moles of solute** in **250 milliliters** of a **1.0 M** solution of $NaCl$?

- A 1.0 mole
- B 0.25 mole
- C 0.50 mole
- D 42 moles



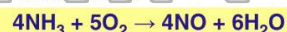
of solution?

- A 1.0 M
- B 2.0 M
- C 0.50 M
- D 0.25 M



10

Given the reaction:



What is the **total number of moles of NO** produced when 1.0 mole of O_2 is completely consumed?

- A 1.0 mole
- B 1.2 moles
- C 0.80 mole
- D 4.0 moles



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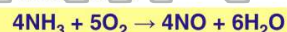
10

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